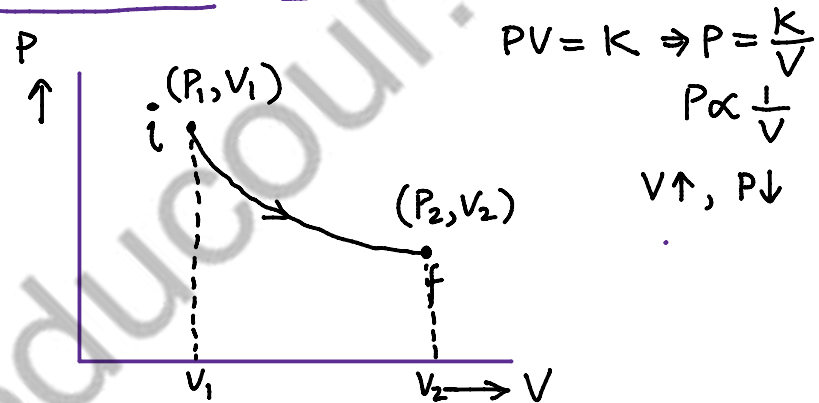


$$T_2 = (200 - 273)^\circ\text{C} = -73^\circ\text{C}$$

Indicator Diagrams for Isothermal and Adiabatic Processes:

(a) Isothermal Process: Isothermal curve



(b) Adiabatic Process: Adiabatic curve

$$PV^\gamma = K \Rightarrow P = \frac{K}{V^\gamma}$$

$$\therefore P \propto \frac{1}{V^\gamma}$$

$$P \uparrow, V \downarrow$$

